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# Mineral Dissolution on Geological Carbon Storage in Carbonates: A Microcalorimetric Study

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Carbon capture and storage (CCS) is a key technology to reduce CO<sub>2</sub> emissions and reach long term climate goals, aiming to limit the temperature rise to 1.5 °C above pre-industrial levels. CCS consists of capturing CO<sub>2</sub> from large industrial processes or from burning fossil fuels in power generation. The captured CO<sub>2</sub> is thereafter transported via pipelines or ships and stored in appropriate geological formations, such as depleted oil and gas reservoirs, unminable coal beds, and deep saline aquifers [1]. The existence of infrastructure, the wealth of reservoir data, and revenue from incremental oil recovery make depleted oil and gas reservoirs the best option for underground CO<sub>2</sub> storage [2]. Carbonate reservoirs may be candidates for CO<sub>2</sub> sequestration through CO<sub>2</sub>-EOR since those reservoirs hold more than 50% of the known petroleum reserves worldwide [3]. The main issue with CO<sub>2</sub> injection into carbonate reservoirs is the formation of carbonic acid that ionizes to form hydrogen ions and bicarbonate ions. This weak acidic environment could alter the performance of carbonate CO<sub>2</sub> storage reservoirs due to dissolution processes that lead to the formation of highly porous and conductive wormholes [4].

Long-term CO<sub>2</sub> sequestration in carbonate reservoirs requires comprehensive assessments of CO<sub>2</sub>-reservoir fluid-mineral interactions. However, this is not an easy task due to the complexity of the reservoir fluid and rock-forming minerals. In the present work, the interfacial-phenomena at rock-aqueous interface during CO<sub>2</sub> injection in carbonate rocks is evaluated by a highly sensitive microcalorimetry technique called that is Isothermal titration calorimetry (ITC). Cobos et al. [5] reported that accurate adsorption enthalpy values for complex rock-fluid systems can be obtained by microcalorimetry. In the ITC experiments, 100 mg of Edwards limestone powder (< 100 µg) was placed in a reaction vessel and 200µL of North Sea formation brine (NFB) was added to the particles. The titration ampule containing the rock-brine slurry was lowered step by step into the calorimeter and equilibrated for 1 hour at 40 °C. Seven injections of 9.948 µL of 3.5 wt% NaCl brine saturated with CO<sub>2</sub> (BCO<sub>2</sub>) were injected independently into the limestone+NFB system with an interval time of 420 seconds between injections. Fluid-fluid experiments consisted in injecting BCO<sub>2</sub> into NFB. Baseline rock-fluid-fluid and fluid-fluid tests were performed similarly but without CO<sub>2</sub>.

This work shows that dynamic rock–fluid and fluid–fluid interactions take place upon CO<sub>2</sub> injection into carbonate rocks due to composition variation “waves” that alters the equilibrium in the system. The results from the Isothermal Titration Calorimetry (ITC) indicate that the dissolution process due to the formation of a weak acidic environment is driven by entropy. This dissolution process is unfavorable with respect to enthalpy change but thermodynamically favorable with respect to entropy change (increment of cations and hydrogen carbonate in the brine). A large perturbation in the water-water network was observed when BCO<sub>2</sub> was injected into the reservoir fluid. This alteration is a result of the salinity difference and also the presence of CO<sub>2</sub> in the injection fluid.

## Participation

In-Person

## References

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